

QP Code: 121006

Reg. No.....

**First Semester B. Pharm Degree Regular/Supplementary
Examinations September 2022
Human Anatomy and Physiology
(2017 Scheme)**

Time: 3 Hours

Max. Marks: 75

- Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space
- Answer all parts of a single question together • Leave sufficient space between answers
- Draw table/diagrams/flow charts wherever necessary

Essay

(2x10=20)

1. Explain the mechanism and factors that influence coagulation.
2. Define cardiac cycle and explain the principle events.

Short Notes

(7x5=35)

3. Discuss the function of Ribosome and lysosomes.
4. Classify joints with examples.
5. Explain epithelial tissue.
6. Discuss structure and function of skin.
7. Explain active transport system across the cell membrane.
8. Muscle contraction.
9. Describe about physiology of hearing.

Answer Briefly

(10x2=20)

10. Define cytology.
11. Functions of mitochondria.
12. Characteristics of connective tissue.
13. Mention the bones of lower limb.
14. Define hypertension and atherosclerosis.
15. Functions of thymus gland.
16. Define parasympathetic and sympathetic nervous system.
17. Homeostasis in biology.
18. Define megaloblastic anemia.
19. Retinoblastoma and glaucoma.

QP Code: 122006

Reg. No.....

**First Semester B. Pharm Degree Regular/Supplementary
Examinations September 2022
Pharmaceutical Analysis - I
(2017 Scheme)**

Time: 3 Hours

Max. Marks: 75

- *Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space*
- *Answer all parts of a single question together • Leave sufficient space between answers*
- *Draw table/diagrams/flow charts wherever necessary*

Essay

(2x10=20)

1. Define different methods of expressing concentration. Write methodology to prepare. • 100 ml of 1N sulphuric acid solution • 250 ml of 0.1N sulphuric acid solution.
2. Write principle involved in non-aqueous determination of weak acids and weak bases. Add a note on pM indicators.

Short Notes

(7x5=35)

3. Explain the principle involved in determination of magnesium sulphate.
4. Explain the method of minimizing errors in pharmaceutical analysis.
5. Write working principle of dropping mercury electrode.
6. With an example, explain the principle of iodometric titration.
7. Classify acid-base titrations. Write one indicator each for different types of acid-base titration.
8. Describe the principle involved in Fajan's method with appropriate example.
9. Write steps involved in estimation of a salt by gravimetric analysis.

Answer Briefly

(10x2=20)

10. Write any two reference electrodes used in potentiometric analysis.
11. Enumerate primary standards used in oxidation-reduction titrations.
12. Define conductance. How is it expressed.
13. Explain amphoteric solvent. Give example.
14. With an example, explain conductometric titration.
15. What are masking and demasking agents in complexometry.
16. What are significant figures.
17. Define reduction. Name any two reducing agents.
18. Primary standards.
19. Explain dichrometry.

QP Code: 123006

Reg. No.....

**First Semester B.Pharm Degree Regular/Supplementary Examination
September 2022**

**Pharmaceutics- I
(2017 Scheme)**

Time: 3 Hours

Max. Marks: 75

- *Answer all questions to the point neatly and legibly* • *Do not leave any blank pages between answers* • *Indicate the question number correctly for the answer in the margin space*
- *Answer all parts of a single question together* • *Leave sufficient space between answers*
- *Draw table/diagrams/flow charts wherever necessary*

Essay

(2x10=20)

1. Describe the history of profession of pharmacy in India in connection with pharmacy education, industry and organization.
2. Define emulsion and write the classification of emulsifying agents with suitable examples. Discuss the methods of preparation, stability problems of emulsions and its rectification

Short Notes

(7x5=35)

3. Evaluation of suppositories.
4. Factors influencing dermal penetration of drugs.
5. Isotonicity and its determination methods.
6. Adult dose of an antibiotic is 500mg. How do you calculate the dose for 10 years' child. (Any two methods)
7. Throat paints.
8. Classification of powders with suitable examples
9. Therapeutic incompatibility with suitable examples

Answer Briefly

(10x2=20)

10. Phase volume ratio
11. Difference between lotions and liniments
12. Displacement value
13. Parts of prescription.
14. HLB Value
15. Preparation of creams.
16. Stokes equation.
17. Types of suspensions.
18. Geometric dilution.
19. Translate the Latin terms in to English

• Hora somni • Omni mane • Stat • Coch mag

QP Code: 124006

Reg. No.....

**First Semester B.Pharm Degree Regular/Supplementary Examinations
September 2022**

**Pharmaceutical Inorganic Chemistry
(2017 Scheme)**

Time: 3 Hours

Max. Marks: 75

- Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space
- Answer all parts of a single question together • Leave sufficient space between answers
- Draw table/diagrams/flow charts wherever necessary • Write the chemical equation/formula wherever necessary.

Essay

(2x10=20)

1. Discuss in detail the apparatus and principle involved in the limit test of arsenic.
2. • Explain the role of electrolyte in acid base balance.
• Give the preparation, assay and uses of sodium chloride.

Short Notes

(7x5=35)

3. What are the antacids. Give the preparations, properties and uses of sodium bicarbonate.
4. What are the properties of α , β and γ rays.
5. What are expectorants. Give the preparation, properties, assay and use of ammonium chloride.
6. Give the principle, and reaction involved in limit test of Iron.
7. Give preparation and properties of sodium iodide I^{131} .
8. Role of fluorides in treatment of dental caries.
9. Explain the methods of adjusting tonicity.

Answer Briefly

(10x2=20)

10. Advantages of combinations of antacids.
11. What are antidotes. Give two examples with formula.
12. Give the principle involved in the limit test of chloride.
13. Write a note on alum.
14. Define isotonicity.
15. What are acidifiers. Give two examples with formula.
16. Give the formula and use of • calcium carbonate • zinc eugenol cement.
17. Properties and use of potassium permanganate.
18. Pharmaceutical application of radioactive substance.
19. Preparation and use of sodium orthophosphate.
