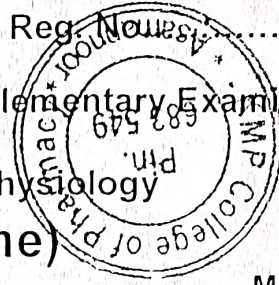


QP Code: 121006

Reg. No.

First Semester B. Pharm Degree Supplementary Examinations
June 2022
Human Anatomy and Physiology
(2017 Scheme)



Time: 3 Hours

Max. Marks: 75

- Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space
- Answer all parts of a single question together • Leave sufficient space between answers
- Draw table/diagrams/flow charts wherever necessary

Essay

(2x10=20)

1. Enumerate the cellular components of blood and their functions. Add a note on iron deficiency of anemia.
2. Explain the neat structure of skull bone, their joints and its functions

Short Notes

(7x5=35)

3. Blood grouping and its significance
4. Circulatory system of heart
5. Draw the structure of human skin with label
6. Scapula
7. Discuss the physiology of cell membrane
8. Mechanism of coagulation
9. Types of synovial joint.

Answer Briefly

(10x2=20)

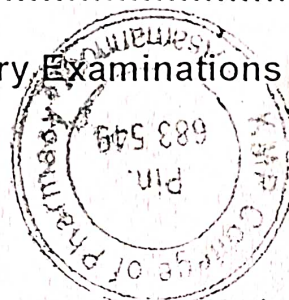
10. Define cardiac output
11. Functions of mitochondria
12. Describe the physiology of vision
13. Auditory ossicles
14. Structure of lymph node and its function
15. Cervical segment of vertebral column
16. Functions of vein
17. Disorders of heart
18. Parts of lower limb
19. Types of leucocytes

QP Code: 122006

Reg. No.....

First Semester B. Pharm Degree Supplementary Examinations
June 2022

Pharmaceutical Analysis - I
(2017 Scheme)



Max. Marks: 75

Time: 3 Hours

- Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space
- Answer all parts of a single question together • Leave sufficient space between answers
- Draw table/diagrams/flow charts wherever necessary

(2x10=20)

Essay

1. What are complexometric titrations. With examples describe of complexometric titrations.
2. Write principle involved in working of glass electrode. Describe typical methods to determine end points in potentiometric titrations.

(7x5=35)

Short Notes

3. Explain the principle involved in determination of sodium chloride by precipitation titration method.
4. Write method of preparation and standardization of 0.1M sodium thiosulphate.
5. Explain types of errors citing appropriate examples in pharmaceutical analysis.
6. Explain the principle involved in conductometric titration of a strong acid with weak base.
7. What is a non-aqueous titration. Explain the principle involved in non-aqueous titration of ephedrine hydrochloride.
8. Describe briefly theory of acid base indicators.
9. With example, write principle involved in diazotization titration.

(10x2=20)

Answer Briefly

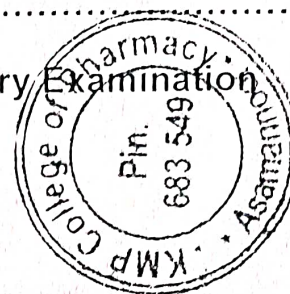
10. Explain secondary standard. Give examples.
11. Define precision.
12. Define molality.
13. What is a protic solvent. Give examples.
14. Explain Ilkovic equation and explain the terms involved in it.
15. Explain post precipitation.
16. Define conductance.
17. What are significant figures.
18. Explain the advantages of cerimetry.
19. Example of acidimetric titration.

QP Code: 123006

Reg. No.....

First Semester B.Pharm Degree Supplementary Examination
June 2022

Pharmaceutics- I
(2017 Scheme)



Time: 3 Hours

Max. Marks: 75

- Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space
- Answer all parts of a single question together • Leave sufficient space between answers
- Draw table/diagrams/flow charts wherever necessary

Essay

(2x10=20)

1. Discuss the stability problems in emulsion. Describe the methods to overcome the stability problems.
2. Give a detailed account on the excipients used in the formulation of liquid dosage forms.

Short Notes

(7x5=35)

3. Classify dosage forms with examples.
4. Differentiate between flocculated and deflocculated suspensions.
5. Advantages and disadvantages of suppositories.
6. Excipients used in semi-solid dosage forms
7. Explain chemical incompatibility with examples
8. Advantages and disadvantages of powders.
9. Preparation of aqueous cream with one example.

Answer Briefly

(10x2=20)

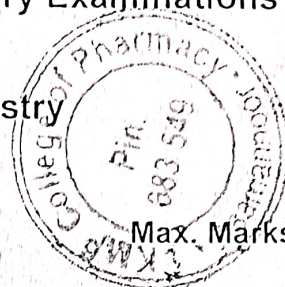
10. Expand the abbreviations and give the English meaning of the following Latin terms: • Pulv • Fiet.mist.mitte
11. Explain idiosyncrasy. Give examples.
12. Define eutectic mixture.
13. Define ointment and give examples
14. Define paste and give examples
15. Define hygroscopic powder and give examples
16. Define suppositories.
17. Define lotion.
18. Therapeutic incompatibility
19. Differentiate between mouthwash and gargle.

QP Code: 124006

Reg. No.....

First Semester B.Pharm Degree Supplementary Examinations
June 2022

Pharmaceutical Inorganic Chemistry
(2017 Scheme)



Time: 3 Hours

Max. Marks: 75

- Answer all questions to the point neatly and legibly • Do not leave any blank pages between answers • Indicate the question number correctly for the answer in the margin space
- Answer all parts of a single question together • Leave sufficient space between answers
- Draw table/diagrams/flow charts wherever necessary

Essay

(2x10=20)

1. • Classify antacids and write a note on ideal properties of antacids and combination of antacids.
 - Give the preparation assay and uses of sodium bicarbonate.
2. • Give the functions of major physiological ions used as electrolytes in the replacement therapy.
 - Give the composition and uses of oral rehydration salt.

Short Notes

(7x5=35)

3. Write about the role of fluorides in the treatment of dental caries.
4. Write a note on Indian pharmacopoeia.
5. Define alum and give the formula properties and use of potash alum.
6. Write the preparation, properties assay and uses of chlorinated lime.
7. Give the properties and use of the following.
 - kaolin
 - sodium ortho phosphate.
8. Write a short notes on iodine and its preparations.
9. Explain the principle with chemical reactions involved in the Limit test of arsenic.

Answer Briefly

(10x2=20)

10. Define haematinics and give one example.
11. Define radioisotope.
12. What is the use of activated charcoal.
13. Principle involved in limit test of Iron.
14. Give two examples of buffer and its application.
15. Properties and uses of sodium iodide – I¹³¹
16. Properties of potassium permanganate.
17. Give the storage conditions of radio isotopes.
18. Properties and use of aluminium hydroxide gel.
19. Write the chemical equation involved and the indicator used in assay of ammonium chloride.
